

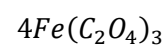
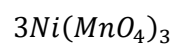
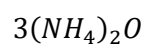
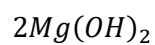
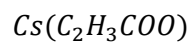
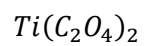
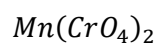
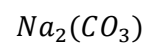
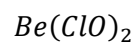
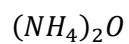
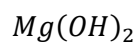
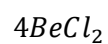
C11 - 1.1 - # of Particles HMW

Fill in the following table.

Element	Atomic #	# of Protons	Atomic Mass	# of Neutrons	# of electrons in <u>Atom</u>	Ion Charge	# of electrons in <u>Ion</u>	Ion Valence electrons
oxygen								
calcium								
silicon								
Iron (II)								
Manganese (IV)								
Nickel (III)								
lithium								
Cobalt (II)								
iodine								
Iridium (IV)								
Lead (II)								
Carbon								
chlorine								
Uranium (VI)								
magnesium								
Chromium (II)								
aluminum								
Platinum (IV)								
helium								
hydrogen								
Plutonium (V)								
Gold (I)								
radium								
neon								
potassium								
tin								
radon								
sodium								
Sulfer								
Mercury								

C11 - 1.1 - # of Atoms WS

How many of each atom are there in the following molecules?



C11 - 1.1 - Bohr/Lewis Diagrams HMW

Draw a Bohr and Lewis diagram for the following atoms.

O

H

Li

Al

N

He

Ne

Fe

K

Na

Be

Draw a Bohr diagram for the following ions

O⁻²

H⁺

Li⁺

Al⁺³

N⁻³

He

Ne

Fe⁺²

K⁺

Na⁺¹

Be⁺²

Fe⁺³

C11 - 1.1 - Bohr/Lewis Diagrams HMW

Draw a Bohr and Lewis diagram for the following Compounds.

O

H

Li

Al

N

He

Ne

Fe

K

Na

Be

Draw a Bohr diagram for the following ions

O⁻²

H⁺

Li⁺

Al⁺³

N⁻³

He

Ne

Fe⁺²

K⁺

Na⁺¹

Be⁺²

Fe⁺³