C11 - 3.2 - Molar Mass WS
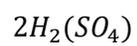
Calculate the molar mass of the following

Cl

H_2

Fe

AlP

K_2O

CO_2

Zn_3P

$H_2(SO_4)$

$NaCl$

MgF_2

$(NH_4)_2O$

$Na_2(CO_3)$

$2Cl$

$3H_2$

$2H_2(SO_4)$

$3CO_2$
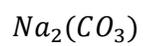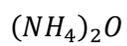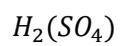
C11 - 3.2 - $g < ->$ Moles WS

How many moles of hydrogen in 1 g of hydrogen?

How many moles of hydrogen in 2 g of hydrogen?

How many moles of nitrogen in 14 g of nitrogen?

How many moles of nitrogen in 42 g of nitrogen?

How many moles of hydrogen in 75 g of hydrogen?

How many moles of NO_2 in 100 g of NO_2?

How many moles of gold in 50 g of gold?

How many moles of Fe_2O_3 in 120 g of iron?

How many moles of calcium in 401 g of calcium?

How many moles of potassium and 3.91 g of potassium?

How many moles of $Cu(SO_4)$ in 32 g of $Cu(SO_4)$?

C11 - 3.2 - Moles$< - > g\ WS$

How many grams of hydrogen in one mole of hydrogen?

How many grams of hydrogen in two moles of hydrogen?

How many grams of nitrogen in three moles of nitrogen?

How many moles of oxygen in 48 g of oxygen?

How many moles of aluminum in 100 g of aluminum?

How many moles of mercury in 200.6 g of mercury?

How many moles of magnesium in 10 g of magnesium?

How many moles of gold in 100 g of gold?

How many moles of helium in 8 g of helium?

How many moles of the nickel in 400 g of nickel?

How many moles of titanium in 47.9 g of titanium?