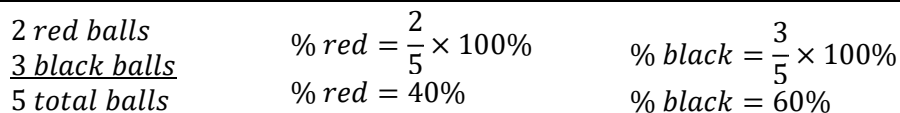
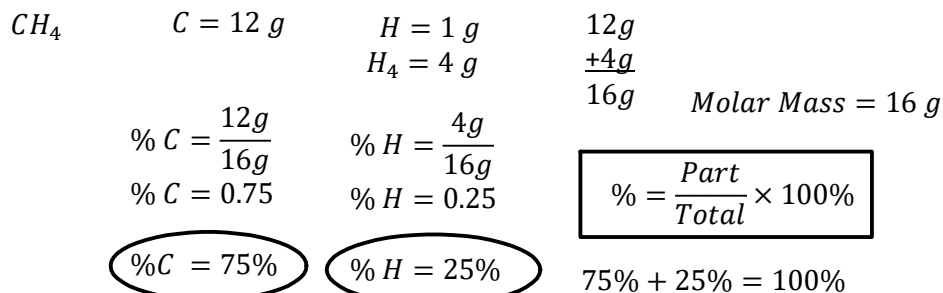


Assume 1 mole

Assume 100g

## C11 - 3.4 - % Composition/Empirical/Molecular Formula Notes

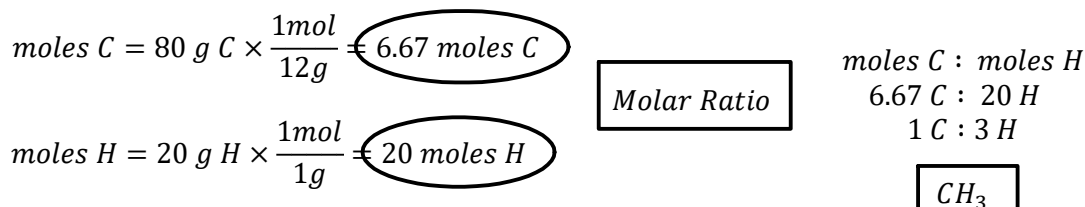
Find the Percent Compotion : Percent by mass of the element in the compound.



Empirical Formula: The simplest formula of the compound.

Assume 100g

What is the Empirical Formula of a compound with 80% Carbon and 20% Hydrogen?



Possibly Double Everything!

Molecular Formula: formula of the compound

Empirical Mass: Molar Mass of the Empirical Formula

A molecule has an Empirical Formula of NO and a Molar Mass of 90 g.

What is the Molecular Formula?

