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STANDARD REDUCTION POTENTIALS OF HALF-CELLS

Ionic Concentrations are at 1 M in Water at 25°C

STRENGTH OF XIDIZING AGENT	OXIDIZING AGENTS		REDUCING AGENTS	Eº (VOLTS)	STRENGTH OF REDUCING AGEN
strong	F ₂ (g) + 2e ⁻	=	2F	+2.87	weak
	$S_2O_8^2 + 2e^-$		2 SO ₄ ²⁻		Weak
A He can be	H ₂ O ₂ + 2H ⁺ + 2e ⁻		2 H ₂ O	+1.78	
E.3	MnO ₄ + 8H + 5e ⁻		Mn ²⁺ + 4 H ₂ O		一名自由 到 自由
AND TO SHALL SHALL SHALL	Au*+ 3e*		Au(s) $\frac{1}{2}$ Br ₂ (l) + 3 H ₂ O		
	$CIO^{-} + 8H^{+} + 8e^{-}$		CI-+4 H ₂ O		
MARIE CONTRACTOR AND	$Cl_2(g) + 2e^{-}$		2 CI	+1.39	
	$Cr_{\circ}O_{\circ}^{2-} + 14H^{+} + 6e^{-}$	=	2 Cr ³⁺ + 7 H ₂ O	+1.23	telow districtly
F. P. SWO	$\frac{1}{2}$ O ₂ (a) + 2H ⁺ + 2e ⁻	=	H ₂ O	+1.23	suki i hedi
	$MnO_2(s) + 4H^+ + 2e^-$	=	Mn ²⁺ + 2 H ₂ O	+1 22	1 2 3 3 3
The state of the s	$10^{-}_{3} + 6H^{+} + 5e^{-}$	-	$\frac{1}{2}I_2(s) + 3H_2O$	+1.20 Overpo	otential
S & CONTRACTOR	Br ₂ (I) + 2e ⁻	=	2 Br-	+1.09 Effec	ct
			Au(s) + 4 Cl-	+1.00	
T. L.			NO(g) + 2 H ₂ O		JUSCOD KO BROOMS
The same of the sa			Hg(I)		
$\frac{1}{2}O_{2}($	$(g) + 2 H^{+}(10^{-7} M) + 2 e^{-}$	=	H ₂ O	+0.82	
100	$2NO_3^- + 4H^+ + 2e^-$	=	N ₂ O ₄ + 2 H ₂ O	+0.80	
MA NEW MENTS	$Ag^+ + e^-$	-	Ag(s)	+0.80	
TE STATE STATE	$\frac{1}{2}$ Hg ₂ ²⁺ + e ⁻	-	Hg(I)	+0.80	auto lain annies
The state of the s	Fe ³⁺ + e ⁻	=	Fe ²⁺		
	$O_2(g) + 2H^+ + 2e^-$	-	H ₂ O ₂		
305			MnO ₂ (s) + 4 OH ⁻		Use to lunci
	$l_2(s) + 2e^-$				
The State of			Cu(s)		
The same town	H ₂ SO ₃ + 4H ⁺ + 4e ⁻	-	S(s) + 3 H ₂ O	+0.45	highest density
	Cu ²⁺ + 2e ⁻	=	Cu(s)	+0.34	
A SERIES	$SO_4^{2-} + 4H^+ + 2e^-$	=	H ₂ SO ₃ + H ₂ O	+0.17	Thyrici dilitalon
Jarra Barra	Cu ²⁺ + e ⁻	=	Cu ⁺	+0.15	The state of the s
THE PERSON NAMED IN	Sn ⁴⁺ + 2e ⁻	=	Sn ²⁺	+0.15	The State of the S
le be	S(s) + 2H ⁺ + 2e ⁻	=	H ₂ S(g)	+0.14	come egod
	2H ⁺ + 2e ⁻	=	H ₂ (g)	0.00	
	Pb ²⁺ + 2e ⁻	=	Pb(s)	-0.13	
	Sn ²⁺ + 2e ⁻	=	Sn(s)	-0.14	
NAME OF STREET	Ni ²⁺ + 2e ⁻				
Nor Espain on co	$H_3PO_4 + 2H^+ + 2e^-$	=	H ₃ PO ₃ + H ₂ O	-0.28	
No. of Persons Reported	Co ²⁺ + 2e ⁻	=	Co(s)	-0.28	
	Se(s) + 2H ⁺ + 2e ⁻	=	H ₂ Se	-0.40	
	Cr ³⁺ + e ⁻	=	Cr ²⁺	-0.41	
			H ₂ + 2 OH ⁻ (10 ⁻⁷ M)	-0.41	
	Fe ²⁺ + 2e ⁻			-0.45	
Overpote	1920(0) . 20	-	2 Ag(s) + S ²⁻	-0.69	
Eff	$Cr^{3+} + 3e^{-}$	=	Cr(s)	-0.74	1 3 4 1 1 1 1 1 1
	Zn ²⁺ + 2e ⁻ Te(s) + 2H ⁺ + 2e ⁻	=	Zn(s)	-0.76	
	Te(s) + 2H ⁺ + 2e ⁻	=	H ₂ Te	-0.79	
The open are a second	2 H ₂ O + 2e ⁻	=	$H_2(g) + 2 OH^-$	-0.83	
The Sales of the sales	Mn ²⁺ + 2e ⁻	=	Mn(s)	-1.19	
	Al ³⁺ + 3e ⁻	=	Al(s)	-1.66	
	$Mg^{2+} + 2e^{-}$	=	Mg(s)	-2.37	
	Na ⁺ + e ⁻	=	Na(s)	-2.71	
THE RESERVE OF THE PARTY OF THE	Ca ²⁺ + 2e ⁻	=	Ca(s)	-2.87	
	Sr ²⁺ + 2e ⁻		Sr(s)	-2.89	
	Ba ²⁺ + 2e ⁻			-2.91	
	K ⁺ + e ⁻			-2.93	1
			Rb(s)	-2.98	V
		=		-3.03	strong
weak	Li ⁺ + e ⁻			-3.04	